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## Closed Cycle Photocatalytic Process for Decomposition of Hydrogen Sulfide to its Constituent Elements DIV

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**Linkous et al.**

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(54) **CLOSED CYCLE PHOTOCATALYTIC  
PROCESS FOR DECOMPOSITION OF  
HYDROGEN SULFIDE TO ITS  
CONSTITUENT ELEMENTS**

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#### Related U.S. Application Data

- (62) Division of application No. 09/375,967, filed on Aug. 17, 1999, now Pat. No. 6,248,218.  
(60) Provisional application No. 60/126,036, filed on Mar. 25, 1999.  
(51) **Int. Cl.<sup>7</sup>** ..... **B01J 19/08**  
(52) **U.S. Cl.** ..... **422/186.03**  
(58) **Field of Search** ..... 422/186.3

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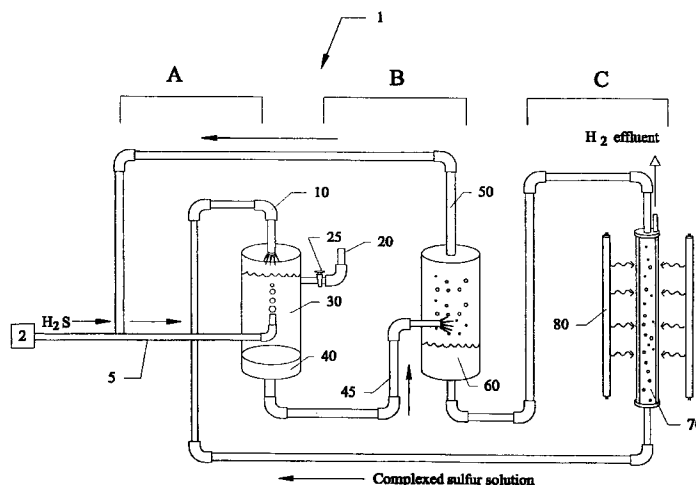
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(57) **ABSTRACT**

System for separating hydrogen and sulfur from hydrogen sulfide (H<sub>2</sub>S) gas produced from oil and gas waste streams. Hydrogen sulfide (H<sub>2</sub>S) gas is passed into a scrubber and filtration unit where it encounters polysulfide solution. Elemental sulfur is freed and filtered through a porous media and continues to a stripper where excess H<sub>2</sub>S is removed. The excess H<sub>2</sub>S returns to the scrubber and filtration unit, while the sulfide solution passes into a photoreactor. The sulfide solution inside the photoreactor is oxidatively converted to elemental sulfur and complexed with excess sulfide ion to make polysulfide ion, while water is reduced to hydrogen. Hydrogen percolates out of the photoreactor, while the polysulfide solution is fed back to the scrubber where the system starts over.

**8 Claims, 3 Drawing Sheets**



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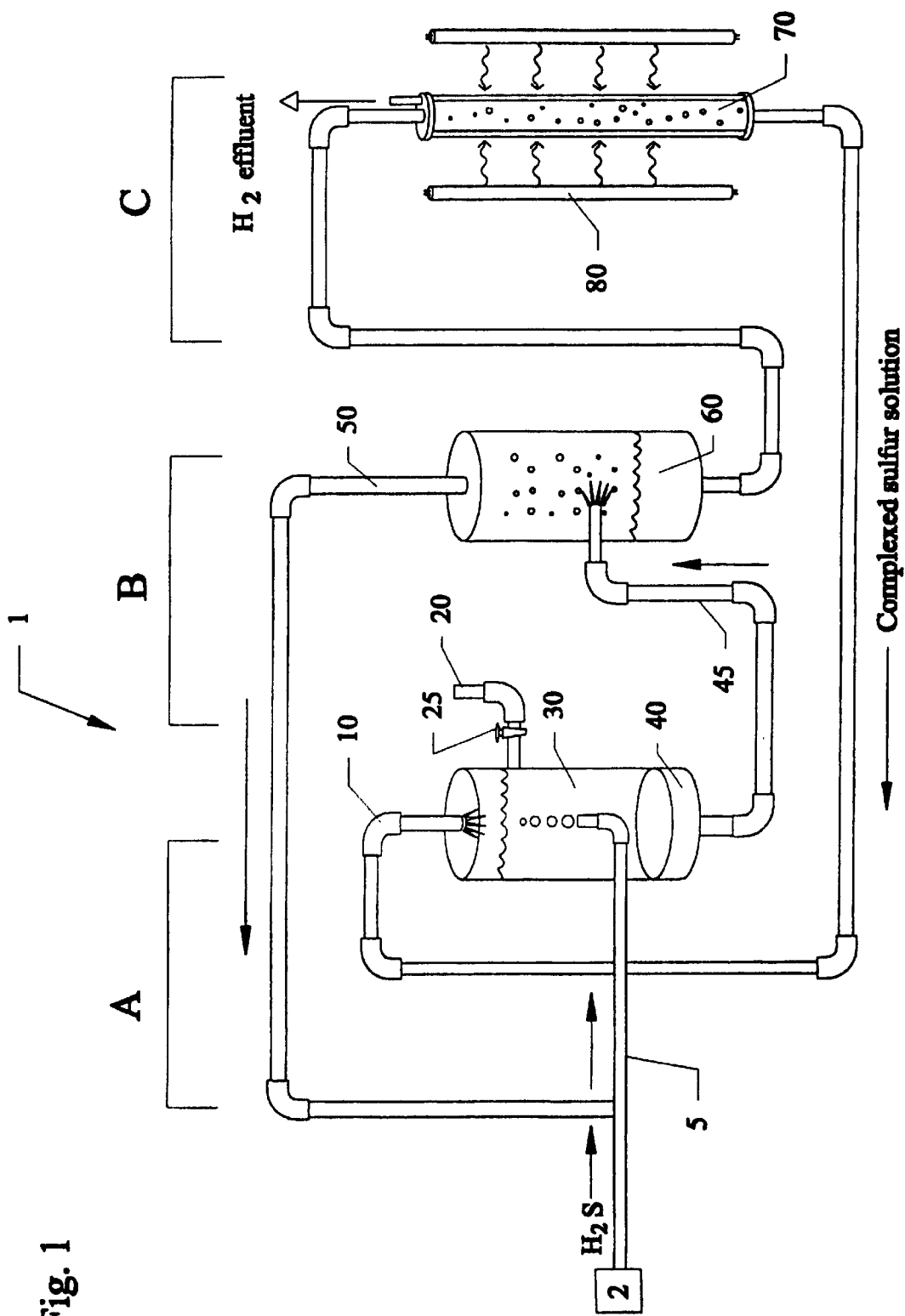


Fig. 1

Fig. 2

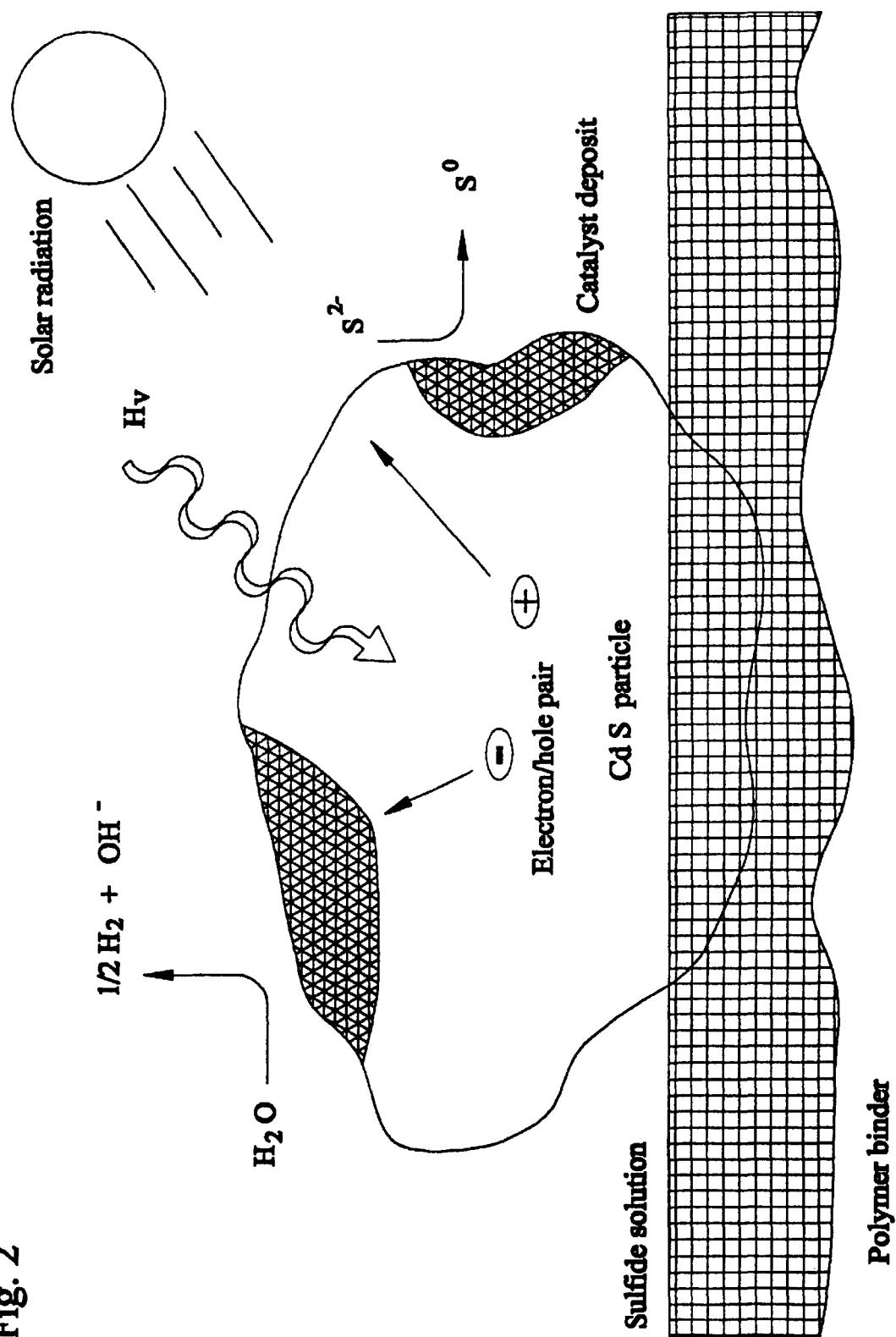
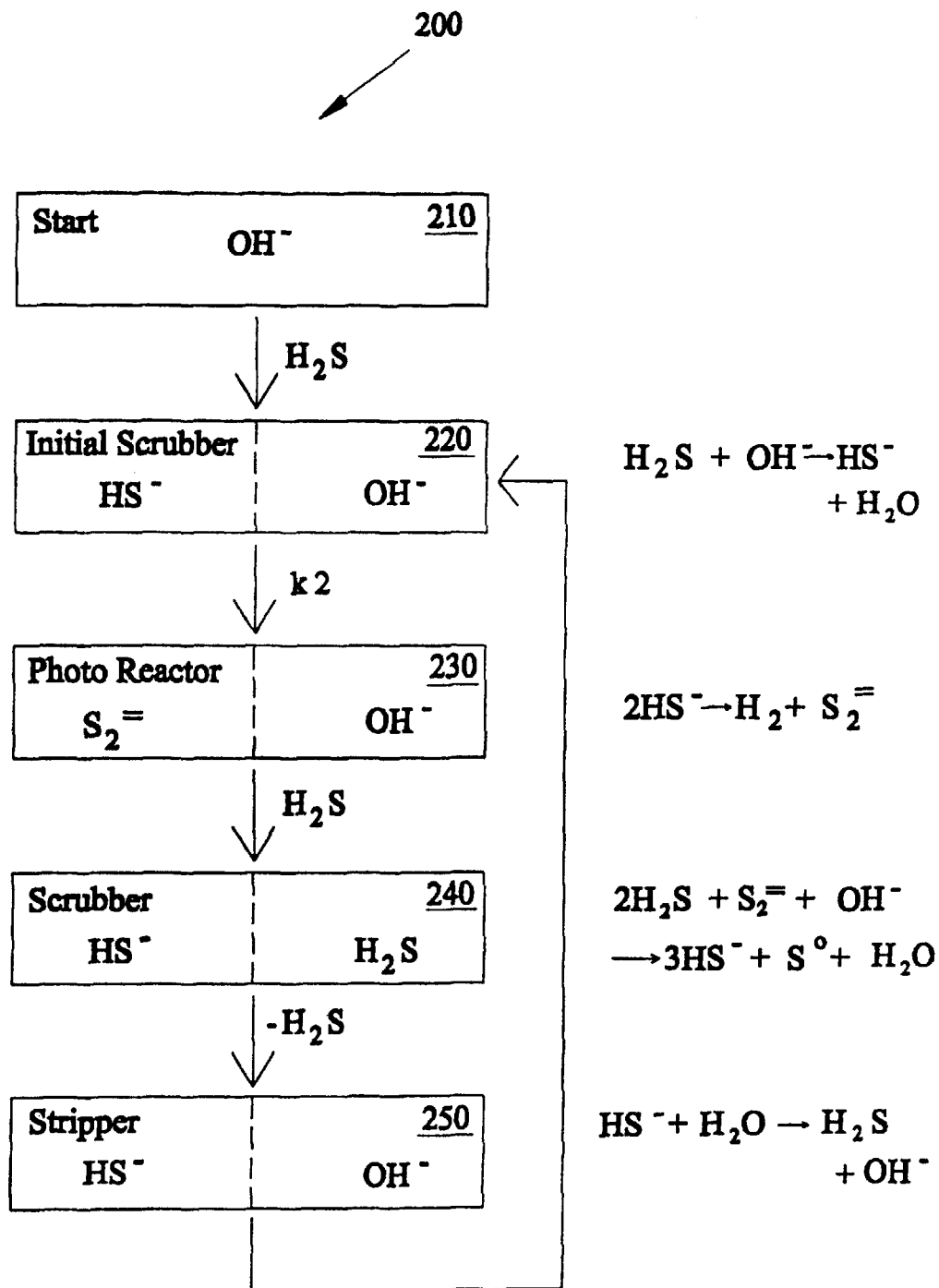


Fig. 3



# CLOSED CYCLE PHOTOCATALYTIC PROCESS FOR DECOMPOSITION OF HYDROGEN SULFIDE TO ITS CONSTITUENT ELEMENTS

This is a Divisional of Application Ser. No. 09/375,967 filed Aug. 17, 1999, now U.S. Pat. No. 6,248,268.

This invention relates to sulfur recovery from oil and gas refinery waste streams, and in particular to a method and apparatus for sulfur recovery from oil and gas refinery waste streams using semiconductor particulates, where hydrogen sulfide ( $H_2S$ ) having been absorbed into an alkaline solution is decomposed in a photochemical scheme to yield both sulfur and hydrogen which can be later sold or used internally at a refinery and is a continuation of Provisional Application No. 60/126,036 filed on Mar. 25, 1999, by the same inventors thereof.

## BACKGROUND AND PRIOR ART

Oil refineries typically pump crude oil from the ground which contains unusable organic sulfur, the latter of which must be separated from the crude oil in order to allow the oil to become a useful product. Typically, the pumped crude oil is treated with hydrogen and a catalyst under a high temperature which decomposes the organic sulfur in the crude oil and releases the sulfur in the form of hydrogen sulfide ( $H_2S$ ) gas. The problem is what do the refineries do with the  $H_2S$  gas, which is a noxious poisonous gas that cannot be released into the atmosphere.

Recent environmental regulations generally require oil and gas refineries to develop waste minimization technologies for the treatment of sulfur-containing compounds, such as hydrogen sulfide ( $H_2S$ ) within their facilities. Current technology typically requires a solvent absorption and stripping process to separate the acid gases, a Claus type reactor to perform a partial oxidation of the sulfide to make elemental sulfur and water and a tertiary gas treatment. The latter of which having specifics that can vary greatly but must effect at least another hundred-fold reduction in  $H_2S$  concentration.

Currently refineries must now go out and buy or make hydrogen to perform hydrodesulfurization. Afterwards, the resultant  $H_2S$  must then be decomposed at the plant.

As the world supply of petroleum decreases, the average sulfur content of crude oil has begun to rise. In recent years the average oil refinery has gone from being a net producer to a net consumer of  $H_2$ . The oil producers would thus be interested in any process that increases the amount of  $H_2$  available in-house.

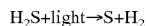
## SUMMARY OF THE INVENTION

A primary objective of the present invention is to provide a method of recovering sulfur from oil and gas refinery waste streams using semiconductor particulates, where hydrogen sulfide ( $H_2S$ ) having been absorbed into an alkaline solution is decomposed in a photochemical process to yield both sulfur and hydrogen.

The secondary objective of the present invention is to decompose noxious poisonous gas such as hydrogen sulfide into separate amounts of hydrogen and sulfur using light and a photocatalyst.

A third objective is to recoup the  $H_2$  contained in  $H_2S$  and return it to the hydrodesulfurization plant for reuse in oil refinery applications.

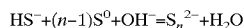
A preferred embodiment of the novel process decomposes  $H_2S$  in a thermodynamically uphill process to its constituent elements:



The energy to drive the reaction can either come from the sun or an artificial light source. The decomposition process is photoelectrochemical in nature, in which the photovoltaic effect from semiconductor solid state physics is combined with conventional electrochemical principles to produce a light-driven electrochemical reactor using semiconductor electrodes. Thus, we call it the PEP (photoelectrochemical particle) process.

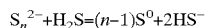
In the preferred embodiment,  $H_2S$  from the stripper is passed into an alkaline solution to produce bisulfide ion,  $HS^-$  (aq). The sulfide solution is then sent into a photoreactor, which consists of a bed of catalyst-activated semiconductor particles such as Cadmium Sulfide ( $CdS$ ), distributed in a thin, planar or cylindrical array depending on the lamp configuration, and a light source containing a wavelength range of approximately 450 to approximately 500 nm. As the solution is passed over the surface of the photo-energized semiconductor particles, bisulfide ion is oxidized and water is reduced to sulfur and hydrogen, respectively. Other photo-atlysts useful can include as Platinized Cadmium Sulfide,  $Pt-CdS$ , Zinc Sulfide,  $ZnS$ , Zinc Ferrate,  $ZnFe_2O_4$ , and Indium Sulfide,  $In_2S_3$ .

The exit stream from the photoreactor will contain sulfur in the form of polysulfide ion, which results when sulfur is produced in the presence of excess bisulfide ion:



The polysulfide solution is returned to the scrubber tank, where pressurizing with  $H_2S$  causes the sulfur to become separated from solution.

The following chemical process shows the effects of combining sulfur recovery and scrubber units, utilizing the acid properties of  $H_2S$  itself to decompose polysulfide ions and free the elemental sulfur.



Further objects and advantages of this invention will be apparent from the following detailed description of a presently preferred embodiment which is illustrated schematically in the accompanying drawings.

## BRIEF DESCRIPTION OF THE FIGURES

FIG. 1 is a layout of the photoelectrochemical particle (PEP) process for  $H_2S$  decomposition for sulfur recovery.

FIG. 2 shows how a semiconductor particle works for sulfur recovery.

FIG. 3 shows a flowchart of the chemical formula reaction occurring in the layout of FIG. 1.

## DESCRIPTION OF THE PREFERRED EMBODIMENT

Before explaining the disclosed embodiment of the present invention in detail it is to be understood that the invention is not limited in its application to the details of the particular arrangement shown since the invention is capable of other embodiments. Also, the terminology used herein is for the purpose of description and not of limitation.

FIG. 1 is a layout system 1 of the photoelectrochemical particle (PEP) process for  $H_2S$  decomposition for sulfur recovery. The system 1 consists of three major units: A, the scrubber and filtration apparatus; B, an outgassing unit such as an  $H_2S$  stripper, and the like; and C, the photoreactor. The scrubber used would be akin to a wet scrubber used in gas

cleaning applications such as but not limited to the scrubber unit used in U.S. Pat. No. 5,211,923 to Harkness et al., which is incorporated by reference. Hydrogen sulfide  $H_2S$ , coming from a source 2 as a hydrodesulfurization plant, a sour natural gas well and the like, is bubbled into scrubber and filtration apparatus A by line 5 where it encounters polysulfide solution 10 coming from photoreactor 70. An outlet 20 with control valve 25 is in place to maintain an overpressure of  $H_2S$  in the head space and to function as a pressure relief valve. The  $H_2S$  lowers the pH of the scrubber vessel solution 30, to approximately 7.0 to approximately 7.5, whereupon the polysulfide anion decomposes, freeing elemental sulfur,  $S^0$ . The sulfur is collected by filtration 40 of the suspension through a porous material such as ceramic fit, and the like.

The filtrate continues by line 45 into the outgassing unit B, where the excess  $H_2S$  is removed from the sulfide solution. The pH of the solution rises to the 10 to 12 range, depending on the extent of  $H_2S$  removal. The  $H_2S$  is returned through line 50 to the foreline of the scrubber unit A, while the sulfide solution 60 is fed to the photoreactor C containing immobilized photocatalyst. The sulfide solution inside the photoreactor 70, consisting mainly of bisulfide ion,  $HS^-$ , is oxidatively converted to elemental sulfur and then complexed with excess bisulfide ion to make polysulfide ion,  $S_n^{2-}$ , while water is reduced to make  $H_2$ , under the action of light source 80. The photocatalyst can be a platinized cadmium sulfide,  $Pt-CdS$ . See for example N.Buhler, et al., Photochemical Hydrogen Production with Cadmium Sulfide Suspensions, *J. Phys. Chem.*, Vol. 88(1984) pp.3261. Other photocatalysts such as zinc sulfide,  $ZnS$ , zinc ferrate,  $ZnFe_2O_4$ , indium sulfide,  $In_2S_3$ , and the like would also additionally have a photocatalytic effect on  $H_2S$ . See Ueno et al., Silica-Supported  $ZnS$  Cds Mixed Semiconductor Catalysts for Photogeneration of Hydrogen, *J.Phys.Chem.*, 89(1985), pp. 3828; Li et al., Hydrogen Production by  $H_2S$  Photocatalytic Decomposition, *New J. Chem.*, 16(1992) pp. 517; and Savinov et al., Suspensions of Semiconductors With Microheterojunctions—A New Type of Highly Efficient Photocatalyst For Dihydrogen Production From Solution of Hydrogen Sulfide and Sulfide Ions, *Int. J. Hydrogen Energy*, 14(1989) pp. 1.

Any broadband visible light source such as a low pressure mercury fluorescent lamp or a tungsten filament bulb, would provide some excitation energy for the photocatalytic reaction. Light sources concentrating their irradiance between approximately 450 to approximately 500 nm would be most effective. This is due to the spectral window created by the absorption edge of the  $CdS$  photocatalyst on the low energy side and the absorption edge of the background polysulfide solution on the high energy side. Examples of such light sources would be  $Zn$ , and  $Cd$  spectral lamps and certain special blue fluorescent lamps.

Referring to FIG. 1, the  $H_2$  gas percolates out the top of the photoreactor C, while the polysulfide solution is returned to the scrubber unit B.  $H_2$  is released out of the top of the photoreactor, underneath "C". The sulfur is shown as being collected by filtration. Use of a settling tank would also be an alternative. Either way, the sulfur is a solid and would have to be collected by opening up the system. Additionally, the sulfur can be melted and collected in molten state.

It would be important to include excess base to keep pH high and facilitate photocatalytic reaction, since the photo-oxidation of  $HS^-$  requires  $OH^-$  as a co-reactant. As previously described a light source having maximum wavelength range of between approximately 450 to approximately 500 nm is needed to penetrate the disulfide solution while remaining within the spectral envelope of  $CdS$ . Use of a reasonably strong acid such as  $HCl$ ,  $H_2SO_4$ ,  $H_3PO_4$  can be used to precipitate sulfur from polysulfide solution. From a

material supply and cost point of view, it would be desirable to use  $H_2S$  (hydrogen sulfide) as the acidic agent.

FIG. 2 shows how a semiconductor particle works for sulfur recovery. FIG. 2 shows the operation of a photocatalytic particle as it would apply to hydrogen sulfide decomposition. The photoparticle is immobilized in a polymer binder and subjected to an ultraviolet (UV) visible radiation source such as sunlight, and the like. Absorption of a photon of light whose energy is greater than the band gap energy of the photocatalytic material results in the creation of an electron/hole pair. The pair represents positive and negative charge carriers that can move independently of each other. Both carriers can be capable of charge transfer into the surrounding solution. The free electron represents reducing, or electron-donating power, while the hole represents oxidizing, or electron accepting, power. Because of the small particle size (on the order of approximately 1  $\mu m$ ) the charge carriers can quickly migrate to the particle/electrolyte interface. The electron serves to reduce water to  $H_2$  gas and hydroxide ion, while the hole serves to oxidize sulfide or bisulfide ion, depending on the solution pH, to elemental sulfur.

FIG. 3 shows a flowchart of the chemical formula reaction occurring in FIG. 1. The process chemistry associate with each part of the PEP process is shown in FIG. 3. For the system start-up, the scrubber contains only caustic solution, representing some background  $OH^-$  concentration (210). Hydrogen sulfide is introduced, generating  $HS^-$  and eliminating half of the  $OH^-$  ion (220). This solution then goes to the photoreactor, where light absorption by the photocatalyst converts bisulfide ion to bisulfide ion,  $S_2^{2-}$  (230). The resultant solution returns to the scrubber, where pressurization by  $H_2S$  results in the reverse conversion plus the precipitation of elemental sulfur,  $S^0$  (240). Stripping away the excess  $H_2S$  raises pH and regenerates the initial balance of bisulfide and hydroxide ion (250).

While the invention has been described, disclosed, illustrated and shown in various terms of certain embodiments or modifications which it has presumed in practice, the scope of the invention is not intended to be, nor should it be deemed to be, limited thereby and such other modifications or embodiments as may be suggested by the teachings herein are particularly reserved especially as they fall within the breadth and scope of the claims here appended.

What is claimed is:

1. A system for recovering sulfur from hydrogen sulfide ( $H_2S$ ) waste streams using a photochemical process with semiconductor particulates, comprising:

- means for producing hydrogen sulfide gas;
- means for mixing the hydrogen sulfide gas with a polysulfide solution to form a mixture solution;
- means for filtering the mixture solution to produce a sulfide solution;
- a stripper means for outgassing hydrogen sulfur from the sulfide solution; and
- a photoreactor for receiving the sulfide solution, the photoreactor having semiconductor particulates and a light source and configured to separate the sulfide solution into hydrogen gas and elemental sulfur.

2. The system for recovering sulfur of claim 1, wherein the means for producing hydrogen sulfide gas is chosen from at least one of:

- an oil refinery, a gas refinery, a hydrodesulfurization plant, and a sour natural gas well.

3. The system for recovering sulfur of claim 1, wherein the photoreactor includes:

- a photocatalyst composed of Cadmium Sulfide ( $CdS$ ).

4. The system for recovering sulfur of claim 1, wherein the photocatalyst includes:



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Platinized Cadmium Sulfide (Pt—CdS).

5. The system for recovering sulfur of claim 1, wherein the photocatalyst includes:

Zinc Sulfide (ZnS).

6. The system for recovering sulfur of claim 1, wherein the photocatalyst includes:

Zinc Ferrate ( $\text{ZnFe}_2\text{O}_4$ ).

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7. The system for recovering sulfur of claim 1, wherein the photocatalyst includes:  
Indium Sulfide ( $\text{In}_2\text{S}_3$ ).

8. The system for recovering sulfur of claim 1, wherein: the light source has a wavelength of approximately 450 to approximately 500 nm.

\* \* \* \* \*