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Apparatus and Method for Photocatalytic Conditioning of Flue Gas Fly-Ash Particles

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A process is provided for conditioning fly-ash present in flue gas which also contains sulfur dioxide and steam by subjecting the fly-ash containing flue gas to ultraviolet radiation so as to convert steam present in the flue gas into hydroxyl radicals which oxidizes about 5 to 10 percent of the sulfur dioxide in the flue gas into sulfur trioxide. This photo-produced sulfur trioxide conditions the fly-ash and renders it more easily separated in the electrostatic precipitator.

3 Claims, 9 Drawing Sheets
OTHER PUBLICATIONS


Fig. 1A

Percent sulfur dioxide conversion

350 ppmv of input SO$_2$, 17 s residence time

- no H$_2$O & no ash, tubular flow
- no H$_2$O & ash deposited on glass wool, tubular flow
- no H$_2$O & compact ash, tubular flow
- H$_2$O & compact ash, tubular flow
- H$_2$O & ash deposited on photoreactor wall, tubular flow
- H$_2$O & ash deposited on photoreactor wall, tubular flow
- H$_2$O & ash deposited on photoreactor inner wall, annular flow
- H$_2$O & ash deposited on photoreactor inner wall, annular flow
Fig. 1B
Circuit diagram for turning UV lamps on/off using an inexpensive operational amplifier and a relay.
Operational logic of the photo catalytic ash conditioning system
Several U.S. patents have been concerned with the ultraviolet treatment of flue-gases. For example, the photolytic oxidation of sulfur dioxide to sulfur trioxide, aided by UV radiation is demonstrated in U.S. Pat. No. 3,984,296 to Richards; 4,097,349 to Zenty; and 5,138,175 to Kim et al. The patents above teach flue gas treatment by a well-known photochemical process involving UV radiation to generate highly oxidizing species that attack target molecules in the flue gas. They may differ on the postulated reaction mechanisms, but they all claim to be processes that remove or help remove gaseous pollutants from the flue gases.

However, the patents discussed above describe methods based on hydroxyl radical reaction engineering and photo-system design. None of these prior art references describe the importance of using water or water vapor solutions in a treatment process and most importantly and fundamentally, they do not claim interactions with fly ash.

U.S. Pat. No. 4,097,349 to Zenty describes oxidation of NOx, SO2, and hydrocarbons with UV radiation having a wavelength of from 240 nanometers to 340 nanometers. Equation 17 of the Zenty U.S. Pat. No. 4,097,349 depicts the absorption of 290–340 nanometers UV radiation to generate a singlet S02. Equation 18 of the Zenty patent describes the absorption of 340–400 nanometer UV radiation by SO2 to produce triplet S02. According to Zenty, through a series of steps, the singlet excited S02 can be transformed to the triplet state, S02. The excited triplet state can be chemically quenched with another species present in the gas stream such as nitrogen, oxygen, water, carbon dioxide, carbon monoxide, ozone, methane, and other hydrocarbons. Zenty’s patent is essentially based on the aforementioned photo-processes and their consequences in photo-oxidation of NOx and SO2. The Zenty patent does not describe any specific mechanisms such as: free radical chain reactions, in general, and hydroxyl radical formation, in particular. For example, Zenty discusses the importance of hydroxyl radical (OH) reactions (column 2, lines 65–68). In the presence of moisture, hydroxyl radical reactions dominate SO2 conversion, as shown in equations 13–16 of the Zenty patent. Fundamentally, the Zenty process does not deal with fly ash conditioning.

U.S. Pat. No. 3,984,296 to Richards describes a process for the reduction of sulfur and nitrogen oxide contaminants in effluent gas streams. Richards’ patent teaches the formation of electron donor-acceptor molecular complexes (EDA complexes) in the flue-gas by exposure to Lewis acids or bases generated electrostatically within a corona precipitator. Also, Richards describes a photochemical technique for the production of the EDA complexes using infrared radiation of 400 to 1,000 nanometers or UV radiation of 120–240 nanometer wavelengths.

Furthermore, Richards describes a technique for photo-induced oxidation of the EDA complexes and reaction of SO2 and NOx molecules with EDA constituent of stack gas. Richards describes using UV light having a wavelength of 150–500 nanometers between 300–400 nanometers to promote photo-oxidation of EDA complex. Richards suggests that free radical reactions may occur due to UV exposure (column 8, lines 12–15). A careful examination of the Richards patent reveals that (Table of column 9, lines 1–21) the underlying reaction mechanisms required for practicing his patent are similarly limited to those disclosed by Zenty as reactions 17–19 of U.S. Pat. No. 4,097,349. Again, Richards does not deal with fly ash.

U.S. Pat. No. 5,138,175 to Kim, et al. describes irradiation of gas mixtures such as combustion gases and flue-gases to facilitate removal of sulfur and nitrogen oxide contaminants.
Kim et al. demonstrates SO₂ can be efficiently removed from flue-gases given sufficient exposure to UV light and presence of adequate amounts of oxygen and water. As far as the photo chemistry is concerned, this patent describes a method for the reduction of sulfur and nitrogen oxides in a gas mixture through UV radiation induced generation of ground state (zero charge atomic) oxygen and subsequent attack of such ground state O₂ upon SO₂ and NO₂. Kim et al., further discloses that a source of radiation having a wavelength of, most desirably, below 220 nanometers installed within a dust-occluding air pressure window device located within the flue-gas stream. What is claimed by Kim, et al., is essentially a sheath design useful for in-situ treatment of flue-gas SO₂ and NO₂. The device supposedly extends the operating life of the lamp and to protect the surfaces of the lamp from fouling.

The patents above teach flue gas treatment by a known photochemical process involving UV radiation to generate highly oxidizing species that attack target molecules in the flue gas. They may differ on the postulated reaction mechanisms, but they all claim to be processes that remove or help remove gaseous pollutants from the flue gases. More importantly, none of the patents discussed above methods and apparatus based on hydroxyl radical reaction engineering and photo-system design.

Several U.S. patents involve in-situ flue-gas fly-ash conditioning and involve techniques for sulfur trioxide injection. U.S. Pat. Nos.: 3,993,429 to Archer; 4,333,746 to Southam; 5,320,052 to Spokoyny et al.; 5,350,441 and 5,196,038 to Wright; 5,229,077 to Bell et al. and U.S. Pat. Nos. 4,966,610 and 5,122,162 to Krigmont et al. (1992) involve various applications to control the addition of a reagent based on measurements of the feedstream and rely on sulfur trioxide injection components along with using Electrostatic Precipitators (ESP) components. However, none of these patents provides for the in-situ transformation of chemical species present in fly ash to form sulfur trioxide for conditioning fly ash.

SUMMARY OF THE INVENTION

The first objective of the present invention is to provide a technique for the in-situ transformation of chemical species already present in flue gas to form small amounts of sulfur trioxide (less than approximately 10% conversion) for conditioning fly ash.

The second object of this invention is to provide a technique that forms sulfur trioxide for conditioning fly ash in coal flue gas without artificially injecting sulfur trioxide into the flue gas.

The third object of this invention is to provide a method of altering the surface electrical properties of fly ash particles in order to condition the fly ash without the injection of chemicals.

The fourth object of this invention is to provide selectively spaced and positioned ultraviolet lamps to maximize generation of the vapor-phase free radical oxidizing species and minimizing mass transfer effects and optimum conditioning of fly-ash particles without the injection of chemicals.

The fifth object of this invention is to condition flue gas containing fly-ash in-situ prior to the treating the flue gas in an electrostatic precipitator. A preferred method of in-situ conditioning fly-ash particles produced by a coal burning power plant includes depositing water vapor and other polar compounds on surface of the fly-ash particles in-situ, wherein the deposition alters electrical resistivity of the particles while photocatalytically converting a small percentage of SO₂ in the flue gas to SO₃ by subjecting the flue gas to mercury vapor UV lamps. The SO₃ conversion is in the range of 2% to 15%, and preferably in the range of 5% to 10%. The UV lamps can be connected in parallel. Each UV lamp can have a diameter of approximately, d, a vertical spacing between adjacent UV lamps of approximately d to 2d, and a horizontal spacing between the adjacent UV lamps of approximately d to 3d, where d can be approximately 1/8 inch to 2 inches. Each UV lamp can further include an angle, θ, of approximately 80 degrees. A large scale preferred embodiment allows for conditioning the flue gas prior to feeding the conditioned flue gas into an electrostatic precipitator (ESP), and then to a stack for expelling the ESP treated flue gas. The large scale embodiment further can include a feedback loop for controlling the turning on and turning off of selected UV lamps. The feedback loop includes a power meter for measuring the power output of the ESP, an opacity sensor for measuring the particulate content in the final atmosphere expelled flue gas, and a computer for controlling each of the UV lamps based on these measurements. The computer can further operate a rheostat for controlling the brightness of each of the UV lamps.

The subject invention does not rely on injection of sulfur trioxide. The subject inventors have disclosed an in-situ transformation of chemical species present in the flue gas to form sulfur trioxide and condition the fly-ash particles by altering their surface electrical properties. More specifically, the subject invention is concerned with fly ash conditioning using plural lamps located at specific positions in a specific arrangement and the chemical conversion that would take place.

Further objects and advantages of this invention will be apparent from the following detailed description of a presently preferred embodiment which is illustrated schematically in the accompanying drawings.

BRIEF DESCRIPTION OF THE FIGURES

FIG. 1A is a graph representing the photolytic conditioning of fly-ash and the effect of moisture using a 16 W UV lamp.

FIG. 1B illustrates the laboratory set-up for demonstrating the effect represented in FIG. 1A of the photolytic conditioning of fly-ash and moisture using a 16 UV lamp.

FIG. 1C depicts a photoreactor containing a layer of compacted ash.

FIG. 1D depicts a thin layer of ash deposited on the outer surface of an inner glass rod with a quartz tube.

FIG. 1E depicts glass wool containing dispersed TiO₂ particles within a quartz reaction tube.

FIG. 2A depicts a cross-sectional view of a preferred light source assembly for in-situ flue-gas fly-ash conditioning.

FIG. 2B depicts a side view of a preferred light source assembly of FIG. 2A along arrow B.

FIG. 2C is a perspective view of the minimal components and circuitry to run high-pressure mercury vapor lamps suitable for large-scale fly-ash conditioning applications with a novel feedback operation.

FIG. 2D shows a conventional wiring diagram for constant wattage ballast used in FIGS. 2A–2C.

FIG. 2E shows a conventional wiring diagram for a capacitor used in FIGS. 2A–2C.

FIG. 2F illustrates a circuit for turning the UV lamps of FIG. 2C on and off.
FIG. 3 is a flow chart depicting the operational logic steps of the preferred embodiment of FIG. 2C.

DESCRIPTION OF THE PREFERRED EMBODIMENT

Before explaining the disclosed embodiment of the present invention in detail it is to be understood that the invention is not limited in its application to the details of the particular arrangement shown since the invention is capable of other embodiments. Also, the terminology used herein is for the purpose of description and not of limitation.

The subject invention is directed to condition the fly-ash particles via hydroxyl radical reactions. Fly-ash consists of fine solids entrained in the flue-gas. What is meant by conditioning is changing the surface properties of the fly-ash particles by deposition and nucleation of strongly polar compounds on the surface in sufficient amounts to alter the electrical resistivity of the particle. This invention does not aim at removing and/or scrubbing gaseous pollutants such as NOx or SOx from flue gases, although a small amount of pollutants will also be removed.

The subject invention takes advantage of the photolytic reaction chemistry and, more importantly, the large effect of moisture (water vapor) present in the flue gas on the hydroxyl radical photo-reactions of importance to ash conditioning. It has been found by the inventors that destruction by photolysis is an order of magnitude slower than by OH radical attack. Table 1 depicts, the relative oxidizing power of hydroxyl radical is highest among all trans-halogen oxidants, surpassed only by fluorine.

<table>
<thead>
<tr>
<th>Species</th>
<th>Oxidation potential (volts)</th>
<th>Relative oxidation power (based on Cl = 1)</th>
</tr>
</thead>
<tbody>
<tr>
<td>F</td>
<td>3.06</td>
<td>2.25</td>
</tr>
<tr>
<td>OH*</td>
<td>2.80</td>
<td>2.05</td>
</tr>
<tr>
<td>Atomic oxygen</td>
<td>2.42</td>
<td>1.78</td>
</tr>
<tr>
<td>O3</td>
<td>2.07</td>
<td>1.52</td>
</tr>
<tr>
<td>H2O2</td>
<td>1.77</td>
<td>1.30</td>
</tr>
<tr>
<td>HO2*</td>
<td>1.70</td>
<td>1.25</td>
</tr>
<tr>
<td>Peroxygencarbonate</td>
<td>1.70</td>
<td>1.25</td>
</tr>
<tr>
<td>Hypochlorous acid</td>
<td>1.49</td>
<td>1.10</td>
</tr>
<tr>
<td>Cl</td>
<td>1.36</td>
<td>1.00</td>
</tr>
</tbody>
</table>

The subject invention focuses on the means to increase hydroxyl radical concentration by improved photo-system design and optimization. The OH radical concentration distribution directly tied to the light intensity profile (irradiance distribution) of the UV lamps. Moreover, the presence of water vapor is thought to be primarily responsible for the observed high SOx conversion efficiencies by increasing the OH radical concentration of the flue-gas stream.

An important aspect of this invention is that fly-ash present in the flue gas provides catalytic surface. Small amounts of semiconductor metal oxide surfaces present in fly-ash (e.g. ferric oxide) act in a manner similar to titania, enhancing the conversion efficiency of the UV reactions, resulting in a fast and economical process.

At the Florida Solar Energy Center (FSEC), the inventors have verified, through laboratory-scale tests, the fact that the conversion of sulfur dioxide to sulfur trioxide occurs by UV irradiation. Bench-scale tests, conducted at FSEC, used a sample of fly-ash from the Crystal River power plant near St. Petersburg, Fla. FIG. 1A depicts results from the laboratory experiments conducted at FSEC. FIG. 1A is a graph representing the photolytic conditioning of fly-ash and the effect of moisture using a 16 W UV lamp. A description is given below.

FIG. 1B illustrates the laboratory set-up for demonstrating the effect of moisture using a 16 W UV(ultra violet) lamp. A quartz tube photoreactor 100 having a 12 mm OD(outter diameter), 10 mm ID(inner diameter) and 220 mm length was placed within an enclosed UV light chamber 110 which was connected to a bubbler 104 and flowed to an external ambient location 107 through a flowmeter 106 and to a Gas Chromotograph with flame photometric detector(GC/FPD) 108. Inner surfaces 112 of chamber 110 had UV reflective surfaces such as but not limited to aluminum and the like. UV illumination was provided by one of two 16 W low-pressure mercury lamps (120A, 120B) located within the light chamber 110 connected by leads 122 to an external electrical power supply 124. To simulate the condition of fly-ash in the flue gas, several arrangements of the quartz tube photoreactor 100 were employed. Air, carbon dioxide (CO2), sulfur dioxide(SO2) and water(H2O) are introduced into the quartz tube photoreactor at position 90 for these arrangements.

In the first arrangement, a layer of compacted ash 130 shown in FIG. 1C filled lower half of the photoreactor 100 was used, and the results are depicted by the “compact ash” notation in FIG. 1A. The “compact ash” arrangement had approximately a ten percent (10%) sulfur dioxide conversion.

The second arrangement includes the set-up of FIG. 1B where a very thin layer of ash 140 was deposited on the inner wall of the quartz tube 100, thus the “wall ash” notation of FIG. 1A. In this arrangement, incoming radiation is filtered by the wall ash limiting gas-particle-photon interactions. In an actual case, UV lamps can be placed within exhaust gas ductwork, directly illuminating the gas and entrained particles.

The third configuration of the tube 100 is shown in FIG. 1D. Here, an annular reactor configuration the inner Pyrex glass rod 150 has an OD of 6 mm with rod 150 placed coaxially within outer quartz tube 100. A thin layer of ash was deposited on the outer surface of the inner glass rod 150. FIG. 1E shows TiO2 particles dispersed throughout a piece of glass wool 160 placed within the quartz reaction tube 100. Using glass wool 160 did not prove very useful as extensive UV filtering by the glass wool did occur.

The results of the laboratory tests in FIG. 1A indicated that sulfur dioxide conversions occurred well above 60% (at a residence time of 17 s). From such rate information, it should be possible to achieve 7-10% SO2 conversion with residence time of only a second or so.

Presence of fly-ash can also enhance SO2 conversion and most importantly, the presence of moisture results in much higher conversions. This demonstrates that the reaction mechanism involves hydroxyl radical formation which renders the photocatalytic oxidation much more efficient compared with the externally generated sulfur trioxide as practiced in the present state-of-the-art ash-conditioning systems.

It is clear from the description above and the background information presented that the sulfur trioxide produced is immediately adsorbed by or reacted with the fly-ash and not emitted from the stack.

Although the hydroxyl radical formation and concentration within the photo-system is not affected by the gas stream advection, the flow (velocity) field does affect the
2A. The height, h, has a value of approximately $d$ to $2d$. The embodiment velocity field and the extent of turbulence and recirculation which are positioned horizontally against the incoming between lamps 332 and 333. Length, $l$, can have a value of becomes very disorderly. The alternating vortices shed into between approximately $d$ and $3d$. Arrow particles within the ductwork.

The character of this wake is dependent on the numerical value of the dimensionless speed called the Reynolds number $Re$ of the flow. For large Reynolds numbers, the wake becomes very disorderly. The alternating vortices shed into the downstream forming a pulsating wake behind the cylinder.

By proper arrangement of the lamps, i.e. their number and relative distance to each other, an optimum photo-system having the most uniform light intensity distribution and highest $SO_2$ conversion is obtained.

FIG. 2A depicts a cross-sectional view of a preferred light source assembly for in-situ fly-gas fly-ash conditioning. FIG. 2B depicts a side view of the preferred light source assembly of FIG. 2A along arrow B. Referring to FIGS. 2A–2B, the embodiment 300 has fly-ash 350 with exhaust gases 360 entering an exhaust ductwork 340 from a coal combustor 302 in the direction of arrow B. In-situ conditioning of the fly-ash 350 with $SO_3$ occurs by passing the fly-ash 350 in the ductwork along arrow 310 by Ultra-Violet (UV) light. Individual lamps 331 have lead wires 352, 354 connected to ballast 210 such as but not limited to 120 V, 60 Hz, 0.625 A such as the one shown in FIG. 2D and connected to main power supply 209 (i.e., 120 volt). The UV light comes from the low pressure mercury vapor lamps 330 which are positioned horizontally against the incoming gases 360 and 350. Any number of commercially available low-pressure or medium/high-pressure mercury vapor lamps having approximately one to five foot length, such as not limited to Voltaire’s G36T6 Ultra V Base and GE 40BL lamps or others such as Model #Philips HOK 140/120, HTQ14 or Hanovia 6850A431 lamps. A single UV lamp 331 contains an external reflective coating 380 that can insulate a minimum surface having a diameter, $d$. The lamp 331 has a grounded shield at 370 which is the same as coating 380. Angle $\phi$ is approximately 80 degrees. This angle designates the angular positions on the cylindrical surface in a cross flow at which flow separation occurs.

The lamps are separated from one another by a vertical height, $h$, represented between lamps 332 and 334 of FIG. 2A. The height, $h$, has a value of approximately $d$ to $2d$. The lamps can further be separated by a length of 1, represented between lamps 332 and 333. Length, $l$, can have a value of between approximately $d$ and 3d. Arrow 320 represents the direction of conditioned flue gas having passed through embodiment 300. The unit $d$, can range from approximately ½ inches to approximately 2 inches. The total number of lamps used is a function of the flue gas flow rate, the concentration of the $SO_3$ and the particulate matter and ash particles within the ductwork.

From the laboratory measurements, the inventors have discovered that residence times of few seconds are needed to achieve approximately 7 to approximately 15% conversion of $SO_3$ to $SO_2$ for fly-ash conditioning, using a 16 W low-pressure mercury lamp. For a typical power plant of 500 MW capacity, the gas flow rate is roughly $4,800,000$ m$^3$/hr (cubic meters per hour) or $1358$ m$^3$/s (cubic meters per second). Assuming a residence time of 1 s, requires $1360$ m$^3$ of reaction volume for full conditioning of the flue gas. This corresponds to duct work dimensions of approximately 25’ by 25’ by 75’ volume. The total number of UV lights required depends on the lamp type and quantum efficiency of $SO_3$ formation. Thus, if 25 ppmv of $SO_3$ is needed, approximately 1.3875 mol/s of sulfur trioxide conversion will be needed. For large-scale applications, the use of high-pressure mercury vapor lamps will be more advantageous. With electrical to photon efficiency of about 35% and quantum efficiency of $SO_3$ formation of 30%, roughly 10% of a lamp’s power input will be useful. Assuming an average photon energy of about 500 kJ, approximately 6.61 MW of power will be required to reform 1.3875 mol/s of $SO_3$ in the flue gas in order to accomplish ash conditioning. This corresponds to about 1.32% of the plant output power generated.

FIG. 2C is a perspective view of a preferred embodiment 400 to run high-pressure mercury vapor lamps 331 suitable for large-scale fly-ash conditioning applications using the components of FIGS. 2A–2B in a novel feedback operation. Flue gas 310 containing fly-ash and exhaust gases(such as 350, 360 discussed previously) from a boiler 290 enters via pipe 295 into ash conditioning photo-duct 340 (also described previously in relation to FIGS. 2A–2B). Duct 340 contains UV lamps 331 with a total output power capable of producing $SO_3$ levels needed for proper ash conditioning. FIG. 2D shows a conventional wiring diagram for constant wattage ballast used in FIGS. 2A-2C. FIG. 2E shows a conventional wiring diagram for a capacitor used in FIGS. 2A-2C. FIGS. 2D and 2E represent schematics of a commercially available high-pressure mercury vapor lamp circuit for the Philips model HTQ-14, 4 kW lamp, of which the circuit is nonessential subject matter which is incorporated by reference. Other appropriate UV lamps include Philips model HOK series with output power exceeding 17 kW.

Referring again to FIGS. 2C–2E, conditioned flue gas 320 passes to a conventional electrostatic precipitator (ESP) 410 (such as the ones described in the background section above). A ESP power meter 420 picks up power signals in kW (kilowatts) from ESP 410 to lamp power controller/computer 430. The opacity sensor 480 is a light beam device that detects the amount of transmitted light being blocked by an exhaust plume 475. The opacity sensor 480 can be an off-the-shelf component such as but not limited to a Datatest single or double pass opacity sensors model #1000 MPS, 1000 MPD or 900 RMD. Conditioned gas from ESP 410 passes through lines 460, 465 into an emitting stack 470 and finally dispersed into the atmosphere as a plume 475. Opacity sensor 480 measures and passes signals to the lamp power controller/computer 430. Output signals 432, 485 from the electrostatic precipitator electrodes power sensor 420 and an opacity sensor 480 installed on stack 470 are processed by a computer controller 430 that regulates total output photonic power within the photo-duct as a feedback to meet the requirements of ash conditioning as the process conditions change. The computer 430 used for controlling the UV lamps can be an IEEE 488 Board and an IBM compatible 586, with a National Instrument Systems data acquisition and control board. The controlling can be done by turning on or off the appropriate number of UV lamps 331 which are used to accomplish ash conditioning in the duct work 340. FIG. 2I illustrates an optional circuit 900 for turning on and off UV lamps 331 which includes opacity sensor 480 connected to an OP-AMP voltage divider 910, through resistor 920(approximately 500-1000 ohms, preferably 720 ohms), to a semiconductor switch 930 such as an
NPN transistor and resistor 940(approximately 10–100 ohms, preferably 55 ohms) through a 12VDC, 115VAC, 350 mA relay 950 to control the turning on and off of lamps 331. All lamps 331 are connected in parallel and each require a capacitor 210 FIG. 2D and a constant wattage ballast 250 shown in FIG. 2E. Ballast 250 and capacitor 210 can be an off-the-shelf Philips 4XVG3/HTQ Ballast with 5 µf farad capacitor, and are nonessential subject matter which are incorporated by reference. The lamp power controller/computer 430 can be further utilized to control rheostat type switches for adjusting the brightness and effective fly-ash conditioning effects of the individual UV lamps 331 in the system 400. Optionally the invention can substitute medium/high pressure mercury vapor lamps, such as Model #Philips HOK140/120, HITQ14 or Hanovia 6850A431 lamps have a built-in rheostat type power modulator for varying and controlling the intensity of each of the lamps, instead of the mercury vapor lamps 331 previously discussed.

FIG. 3 shows a flow chart depicting the operational feedback logic steps of the preferred embodiment of FIG. 2C. The system is started 510 and reset 515 on each pass. From the reset 515, the computer 430 of FIG. 2C reads the ESP power consumption 520 to determine if ESP power usage is within preselected ranges 525. If the answer is no the next step is to increase the number of UV lamps that are lit, 540, followed by signaling to UV lamp power controller 430 to turn on the lamps 331, shown by step 544. Next a timer 536 restarts the beginning of the flow chart measuring after a preselected time period 536. If ESP power usage is determined at step 525 to be within a preset range, then step 520 occurs where the opacity meter is read. The next step is determining whether the opacity reading is within a preset range at step 533. If the answer from step 533 is no the next step is 540, if the answer is yes the system passes to step 536, and so forth.

In all flue gas treatment methods the objective is a substantial reduction in the concentration of the flue gas contaminants, most desirably at 90% levels. Thus, for SO₂ treatment, a conversion efficiency of at least 90% would be most desirable. Unlike processes aimed at the treatment of the flue gas constituents, fly-ash conditioning only requires approximately a 5% to 10% SO₂ conversion.

While the invention has been described, disclosed, illustrated and shown in various terms of certain embodiments or modifications which it has presumed in practice, the scope of the invention is not intended to be, nor should it be deemed to be, limited thereby and such other modifications or embodiments as may be suggested by the teachings herein are particularly reserved especially as they fall within the breadth and scope of the claims here appended.

We claim:

1. A method of in-situ conditioning fly-ash particles produced by a coal burning power plant upstream to an electrostatic precipitator without injecting sulfur trioxide, comprising the steps of:
   producing a flue gas containing fly-ash particles and SO₂, the fly-ash particles having catalytic surfaces;
   photocatalytically converting approximately 5% to approximately 10% of the SO₂ to SO₃; and
   conditioning the fly-ash with the photocatalytically produced SO₃.

2. The method of conditioning fly-ash particles of claim 1, wherein the photocatalytic step includes supplying the ultraviolet radiation.

3. The method of conditioning fly-ash particles of claim 2 wherein the ultraviolet lamp further includes supplying the ultraviolet radiation with vertical and horizontal adjacent ultraviolet lamps.

* * * * *
It is certified that error appears in the above-identified patent and that said Letters Patent is hereby corrected as shown below:

Col. 1, lines 13-14 should read,

STATEMENT REGARDING FEDERALLY SPONSORED RESEARCH OR DEVELOPMENT

This subject invention was made with government support under the DOD/Navy/NSWC, federal contract number N0017491C0161. The government has certain rights in this invention.

Signed and Sealed this
Eighth Day of March, 2011

[Signature]

David J. Kappos
Director of the United States Patent and Trademark Office